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Modern Chemistry Chapter 6. Chemical Bonding, Covalent & Molecular Compounds, Ionic Bonding & ionic Compounds, Metallic Bonding, and Molecular Geometry. chemical bond. a mutual electrical attraction between the nuclei and the valence electrons of different atoms that binds the atoms together.

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Covalent bond in which the bonding of electrons are shared equally by the bonded atoms, which results in a balanced distribution of electrical charge (Electronegativity of 0-0.3) Electronegativity. The measurement of the attraction of the atom for a shared pair of electrons.

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Ionization energy is the energy required to remove an electron from a gaseous atom. 8. A high ionization-energy value indicates that the atom has a strong hold on its electrons and is not likely to lose an outer electron and form a positive ion.

Ch 6 Study Guide answers

CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

6 Chemical Bonding - Effingham County Schools / Overview

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Modern Chemistry 47 Chemical Bonding CHAPTER 6 REVIEW Chemical Bonding SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. ____ In metals, the valence electrons are considered to be (a) attached to particular positive ions.

Modern Chemistry Chapter 6 Review Answers Section 3

Modern Chemistry 7 Solutions CHAPTER 12 REVIEW Solutions SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1.

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Describe the errors made by the following students in making molar solutions. a. James needs a 0.600 M solution of KCl. He measures out 0.600 g of KCl and adds 1 L of water to the solid.

CHAPTER 12 REVIEW Solutions

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Modern Chemistry 46 Chapter Test Chapter: Chemical Bonding In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1. The charge on an ion is a. always positive. b. always negative. c. either positive or negative. d. zero. ____ 2.

Assessment Chapter Test A - Kettering City School District

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