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– where k is the rate coefficient, A is the pre-exponential factor, is the activation energy, is the temperature, and is the ideal gas constant. By linearizing the equation, one finds that $\ln(k) = -\frac{E_a}{R} \left(\frac{1}{T}\right) + \ln(A)$ such that plotting $\ln(k)$ vs.

Chemical Engineering Kinetics - Tufts University

The reactor was designed for 6.5 L total and 4.5 L working volume; the power rating and power delivery were 2.0 and 1.5 hp, respectively. The operating conditions for the reactor to attain its highest yields of 91.01% and 76.67% with the two alternative methods had 800 rpm agitation speed and 40°C temperature.

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Chemical kinetics, also known as reaction kinetics, is the branch of physical chemistry that is concerned with understanding the rates of chemical reactions. It is to be contrasted with thermodynamics, which deals with the direction in which a process occurs but in itself tells nothing about its rate.

Chemical kinetics - Wikipedia

Butt, JB & Broadbelt, LJ 2000, Solutions Manual for Reaction Kinetics and Reactor Design. 2nd edition edn, Marcel Dekker, Inc, New York.

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